Name: _____

Date: _____ Grade: _____

SECTION QUIZ 1-2



Remember:
Score an average of 85% or higher on all three quizzes, and you may be exempt from the chapter test!



Congratulations! You have completed this section!
Answer the following questions to test your understanding of the material. When you are finished, hand in your quiz to your teacher. Each answer is worth 2.2 points.

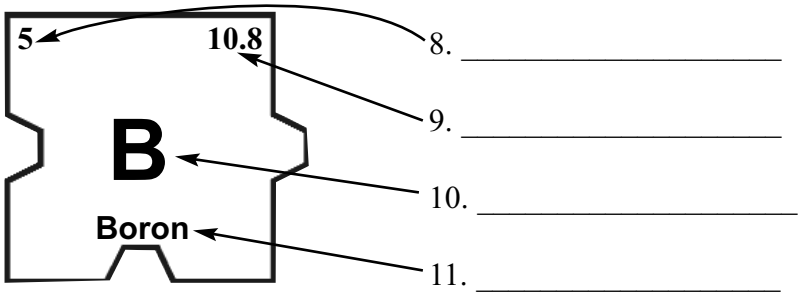
Match the following:

1. _____ Greeks — (a) Atomic Theory: all matter is composed of atoms
2. _____ Democritis — (b) originated the word "atom," meaning "indivisible"
3. _____ John Dalton — (c) matter is made up of small particles with common properties

Matching:

4. _____ Nucleus — (a) discovered by Ernest Rutherford in 1919, has a positive charge
5. _____ Proton — (b) discovered by J. J. Thomson in 1897, has a negative charge
6. _____ Neutron — (c) discovered by James Chadwick in 1932, has a neutral charge
7. _____ Electron — (d) center of an atom, has a positive charge, composed of protons and neutrons

Label the diagram of the element, Boron:



Match the scientists:

12. _____ Ernest Rutherford
13. _____ Neils Bohr
14. _____ Aristotle
15. _____ Jons Jakob Berzelius
16. _____ Dmitri Mendeleev
17. _____ Henry Cavendish

(a) Swedish chemist, proposed using symbols to represent elements
(b) Greek philosopher, believed only 4 elements existed: air, fire, water, earth
(c) British scientist, disproved the 4 basic elements theory about 1780
(d) Russian chemist, developed Periodic Table of the Elements in 1869
(e) Danish scientist, developed circular orbit model of the atom in 1913
(f) British scientist, proposed planetary model of the atom in 1911

SECTION QUIZ 1-2

Fill in the blank. Use words from the WORD BANK.

WORD BANK		
element	solar system	nucleus
pattern	periods	families
weight	atomic activity	compound

18. Together, protons and neutrons compose the _____ of an atom.
19. The most accurate models of the atom resemble the pattern of our _____.
20. Matter made up of just one kind of atom is called an _____.
21. The _____ of protons and neutrons in the nucleus is called the atomic mass of the atom.
22. The chart developed by Mendeleev was based on properties of groups of elements that repeat in a _____.
23. Elements in a vertical column are called _____.
24. Elements in a horizontal row are called _____.

True or False: Write the correct answer.

25. _____ Both atomic mass and atomic numbers are listed on the Periodic Table in order from largest to smallest.
26. _____ All known matter in the universe is made up of elements, either alone or in different combinations.
27. _____ An element is a substance which cannot be broken down into anything simpler by ordinary physical or chemical means.
28. _____ Most elements are made artificially; however, some occur naturally.

SECTION QUIZ 1-2

Matching:

29. _____ Periodic Table of the Elements
30. _____ atomic number
31. _____ atomic mass
32. _____ alkali metals
33. _____ Periodic Table families
34. _____ noble gases
35. _____ atom stability

(a) resemble each other and react similarly, are designated by Roman numerals
(b) most reactive, least stable family group
(c) equals the number of protons in an atom, listed at the top of each element box, always a whole number
(d) depends on number of electrons in outer orbit
(e) includes an atom's neutrons and protons, always expressed as a decimal
(f) least reactive, most stable family group
(g) a chart identifying the elements and their properties
(h) alkaline-Earth metals, family II; less reactive than family I

Use the following table below to answer questions #36-38.

Energy Level	MaximumNumber of Electrons Held
Level One	2
Level Two	8
Level Three	18
Level Four	32

36. The first energy level, closest to the nucleus, can hold a maximum of _____ electrons.
37. Which energy level can hold a maximum of 18 electrons? _____
38. How many electrons can be held on level four? _____

Use the following table below to answer questions #39-41.

Particle	Relative Mass	Charge	Placement
Proton	1	positive	nucleus
Neutron	1	neutral	nucleus
Electron	0.0005	negative	energy level around the nucleus

39. The proton and the _____ have about the same mass.
40. What is the placement of the neutron? _____
41. Which have the same placement? _____
    (a) proton and electron
    (b) proton and neutron
    (c) electron and neutron

SECTION QUIZ 1-2

Complete the Life Principles.

42. *"Some things have to be _____ to be seen."*

43. *"Setting a good example is a _____ for others to follow."*

44. *"A wise person _____ the things in his or her life that need changing and accepts those things that cannot be changed."*

45 *"Things done decently and in _____ improve understanding and efficiency."*

46. *"Basic _____ of love, joy and peace characterize stable, happy families."*